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Toronto, May, 1868.

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The extra quality are equal to Lubin or Rimmel's Perfumes, at 30 per cent. less cost. Hair Oils, Pomades, Tooth Washes, Tooth Powders, Colognes, Lavanders, Sachets, Camphor Ice and Roll, Toilet Vinegar, Milk of Roses, etc., in all the popular styles.

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rolled up and kept in a canister." The coating of gutta percha, through which the acid permeates freely, is given to prevent it adhering to the skin, as "it is desirable that such a dressing should adhere very slightly, if at all. It has this great advantage over the antiseptic lead plaster, that it cannot be softened either by a watery or an oily fluid." If made to contain much less than 25 per cent. of the acid, it is brittle, but this may be avoided by the addition of spirit of wine in an equivalent quantity, as this sample contains 12½ per cent. of acid and the same of spirit.

These plasters are generally kept applied to the part by means of ordinary adhesive plaster strapped round the edges of the piece employed. But to avoid any chance of germs getting access to the wound, to the adhesive plaster before spreading, he directs 1 per cent. of carbolic acid to be added.

Many other applications are used in this system of treatment, but these plasters being interesting pharmaceutical preparations, I have thought worthy of bringing under your notice this evening.

The samples exhibited were prepared in the Hospital Dispensary. — *Pharmaceutical Journal (Eng.)*.

**Hair and Hair Dyes.**

The attention which we called some time since, to the new and perfect black hair dye which Dr. McCall Anderson lately incidentally hit upon, produced a long series of commentaries from accomplished dermatologists and others, well qualified to speak on the not uninteresting subject. Mr. Erasmus Wilson, a leader amongst the professors of dermatology, now enters upon, and discusses the whole question in a series of very interesting observations in the *Journal of Cutaneous Medicine*. He observes, that the hair owes its property of dyeing to its porosity; which is evidently greater than its physiological structure would lead us to infer. Another of its properties, namely, the presence of sulphur in its constitution, renders it prone to darken under the use of certain mineral substances; for example, lead and mercury, whose compounds with sulphur are black. Thus if a weak solution of lead or mercury be brushed into the hair, a certain quantity of the solution will penetrate the hair, and a dark color will be produced in consequence of the formation of a sulphuret of lead or sulphuret of mercury. The depth of the shade of color will depend upon the quantity of sulphur present in the hair, and as red hair and light-colored hair contain more sulphur than dark hair, the result will in that case be comparatively greater. But where the amount of sulphur is too minute to produce the dye, science suggests the means of introducing more sulphur, as is illustrated, by a reversal of the process, in the following quotation from a paper by Dr. McCall Anderson on *Eczema marginatum*: "During the treatment I accidentally discovered what promises to be the most perfect black dye for the hair which has been seen. After having used the bichloride lotion for some weeks, I changed it for the lotion of hyposulphite of soda and the morning after the first application, the hair of the part which before was bright red, had become nearly black. One or two more applications rendered it jet-black, while neither the skin

nor the clothing was stained. I saw this patient a couple of weeks later, and there was not the least deterioration of color; although, of course, as the hair grows the new portions will possess the normal tint." The reason of the escape of the epidermis, while the hair was so thoroughly dyed, is that it contains no sulphur. Mr. Balnanno Squire, in a commentary on the above process, observes that if instead of the hyposulphite of soda one of the more common mordants be employed—say, for example, the sulphide of ammonium, "instead of a black, a bright red color will result. The *modus operandi* of Dr. Anderson's dye is this. The hyposulphurous acid, on being liberated from the soda, decomposes into sulphurous acid and sulphur. The sulphurous acid reduces the bichloride of mercury to the chloride, and the sulphur converts the chloride into (black) sulphide. The effect of the sulphide of ammonium on bichloride of mercury is to produce the (red) bisulphide which is the common vermilion of commerce." Another commentator on "hair dyes" observes that, with the barbers the "sheet-anchor appears to be lead and lime." And again it is recommended to "first wash the hair with a solution (ten grains to the ounce) of nitrate of silver; then use a weak solution of pyrogallie acid, and wash." An interesting article on the subject, from the pen of an able chemical writer, Dr. Scoffern, may be found in the May number of *Belgravia*, under the head of "Cosmetics for the Hair." Dr. Scoffern reminds us that the Persians employ indigo to procure a blue-black dye, and the Turks and Egyptians a "pasty writing ink," composed of pyrogallie acid in combination with a native ore of iron, while in the West the chief constituent of hair-dyes are metallic bodies and walnut-juice. The metals chiefly in use as "capillary chromatics" are silver, lead, and arsenic; while others applicable to a similar purpose are gold, bismuth, iron, copper, cadmium, titanium, uranium, and molybdenum. Lead, in its crudest form, is represented by the leaden comb; but as the process by this means is slow, a compound of oxide of lead or litharge, with lime, and made into a paste with water, is more commonly employed. This is smeared on the hair at night, the evolved gases being imprisoned by an oilskin cap, and in the morning the dried paste is brushed out, and the hair refreshed with pomatum. Or, if a so-called brown, a "smothered" or "fusty black" be required, the paste should be mixed with milk instead of water. The night is preferable for these remedies, because the hair is supposed to exhale more sulphur at this period than during the day. These preparations remind us of a lotion in common use at the present time, consisting of a drachm of acetate of lead with twice the quantity of sulphur to half a pint of water. The nitrate of silver is another common form of dye, but it is open to the objection of staining the skin, and, in fact, everything it touches, and also of becoming iridescent on exposure to light, producing, as Dr. Scoffern observes, a "chromatic play of tints," which is very undesirable. Bismuth presents the same characteristics as lead, but is not much used; and when iron is employed to produce a black tint, it requires for its mordants either the pyrogallie acid or the hydrosulphate of ammonia. Brown is produced by the chloride of gold alone, as also by a solution of sulphate